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provided 1 Calculate the following values without using a calculator 40 a The $[H^+]$ in a solution is $1 \times 10^{-4} M$ Calculate the pH ___; $1 \times 10^{-4} \dots$

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continued 6. Three
students were asked to
determine the volume of
a liquid by a method of
their choosing. Each did

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three trials. The table
below shows the results.

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pH SHORT ANSWER

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40 a The $[H^+]$ in a solution is $1 \times 10^{-4} M$
Calculate the pH

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pH. SHORT ANSWER .

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questions in the space

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provided. 1. Calculate the following values without using a calculator. 4.0 . a. The $[H_3O^+]$ in a solution is 1.0×10^{-4} M. Calculate the pH. _____; 1.0×10^{-11} M. Calculate the pH of a solution is 1.0×10^{-3} M. _____ b. The pH of a solution is 13.0. Calculate the $[H_3O^+]$

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react with active metals
and release hydrogen gas.

- Acids react with bases
to produce $\text{SO}_2 + \text{H}_2\text{O}$

$\text{H}_2\text{SO}_3 \text{ SO}_3 + \text{H}_2\text{O}$

$\text{H}_2\text{SO}_4 \text{ NO} + \text{H}_2\text{O}$

$\text{HNO}_2 \text{ NO}_2 + \text{H}_2\text{O}$

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