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7.B Mole and Mass Avogadro's
Number, The Mole, Grams, Atoms,
Molar Mass Calculations - Introduction
Converting Between Grams and Moles
Mole Ratio Practice Problems Concept
of Mole - Part 1 | Atoms and Molecules
| Don't Memorise Empirical Formula

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\u0026 Molecular Formula

Determination From Percent

Composition Molality Practice

Problems - Molarity, Mass Percent,
and Density of Solution Examples

~~Introduction to Moles~~ Stoichiometry

Basic Introduction, Mole to Mole,

Grams to Grams, Mole Ratio Practice

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Problems Mole Conversions Made
Easy: How to Convert Between Grams
and Moles How To Convert Grams To
Moles - VERY EASY! Counting Atoms:
Intro to Moles Part 2 ~~Converting
Grams to Moles Using Molar Mass |
How to Pass Chemistry The Mole:
Avogadro's Number and Stoichiometry~~

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Introduction to Combustion Analysis,
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Formula Problems Dalton's Law of
Partial Pressure Problems \u0026amp;
Examples - Chemistry ~~Stoichiometry~~
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CHEMISTRY WORKSHEET # 6
MIXED MOLE PROBLEMS (GRAMS,
MOLECULES, AND LITERS) You now
know three things a mole can be: a
molar mass, 6.02×10^{23} molecules
and, for a gas, 22.4 liters at STP.

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Page 10/38

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NAME: HR: _____ PAGE 1

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NAME: _____ HR: _____ PAGE 4

CHEMISTRY WORKSHEET # 2: THE MOLE AS A UNIT OF MASS Now that you know how to find the mass of one mole of a substance (molar mass) you can easily find the mass of several

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moles or the mass of a fraction of a mole using the factor-label technique.

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Sep

A mole of carbon contains 6.02×10^{23} atoms of carbon. A mole of helium contains 6.02×10^{23} atoms of helium.

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A mole of sodium contains 6.02×10^{23} atoms of sodium. A mole of gold contains 6.02×10^{23} atoms of gold. However, we have learned about seven elements that exist as diatomic molecules— H_2 , N_2 , O_2 , F_2 , Cl_2 , Br_2 , and I_2 .

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NAME: _____ HR: _____ PAGE

4CHEMISTRY WORKSHEET # 2:

THE MOLE AS A UNIT OF MASS
Now that you know how to find the mass of one mole of a substance (molar mass)

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you can easily find the mass of several moles or the mass of a fraction of a mole using the factor-label technique.

1 mol = a molar mass of an

atom/molecule (g/mol) EXAMPLE:

What is the mass of 5.00 moles of

water H₂O? STEP 1: 2 H = 2 x (1.01) =

2.02 STEP 2: ? grams H₂O = 5.00

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moles $H_2O \times 18.02 \text{ g } H_2O = 90.10 \dots$

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6CHEMISTRY WORKSHEET # 4

MOLAR MASSES AND

AVOGADRO'S NUMBER Now that

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you know two definitions of a mole (a molar mass and an Avogadro's number of particles) you can combine these two definitions into one problem. EXAMPLE: How many molecules are there in 90.1 grams of water?

$2 \text{ H} = 2 \times (1.01) = 2.02$
 $\text{O} = 1 \times (16.00) = 16.00$
 $\text{H}_2\text{O} = 18.02 \text{ g}$

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molecules $\text{H}_2\text{O} = 90.1 \text{ g } \text{H}_2\text{O} \times 1$
mole $\text{H}_2\text{O} \times 6.02 \times 10^{23}$ molecules
 $\text{H}_2\text{O} = 3.01 \times 10^{24}$ molecules $\text{H}_2 \dots$

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Page 29/38

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NAME: _____ HR: _____ PAGE 9

CHEMISTRY WORKSHEET # 6

MIXED MOLE PROBLEMS (GRAMS,
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know three things a mole can be: a molar mass, 6.02×10^{23} molecules and, for a gas, 22.4 liters at STP. We can use this

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Question: PAGES CHEMISTRY B-
MOLES PACKET NAME: HR:
CHEMISTRY WORKSHEET #5 MOLE
PROBLEMS-MOLAR VOLUME OF A

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GAS We Have Learned That A Mole Is A Mass Of Material And Number Of Particles. A Mole Can Also Be A Measure Of Volume When We Are Talking About Gases. You May Remember From Previous Science Classes That All Gases Behave Basically The Same As Far As The

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Page 6/9

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Page 36/38

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